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below in your answer boyle s law
charles s law kinetic molecular theory
pressure 1 temperature pressure and
volume affect the behavior of gases
which variable is held constant in the
relationship described by boyle s law 2
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web the amount of space that matter
fills measured in cubic centimeters cm³
milliliters ml liters l gas particles move
and fill the space available this
measure of gas is the same as the
measure of its container a measure of
the average energy of random motion
of the particles of a substance
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to fill its container unlike solids or
liquids the reverse is also true they are

easily compressed or squeezed into a
smaller volume compressibility is a
measure of how much the volume of
matter decreases under pressure
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behavior of gases show full text are
fluids have no definite shape or volume
and they expand to fill their container
particles move rapidly in all directions
molecules are in constant motion and
collide with each other and walls of the
container low density because particles
are far apart are compressible

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