

## Léo Pomar

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of those who partake in its reading experience.

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oxide  $i_2O_5$  reacts with 28 0 grams of

carbon monoxide co determine the mass of iodine i<sub>2</sub> which could be produced b if in the above situation only 0.160 moles of iodine i<sub>2</sub> was produced i what mass of iodine was produced

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage

of the theoretical yield  
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web start with 8.3 mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation  $8.3 \text{ mole h}_2\text{s} \times \frac{2 \text{ mole h}_2\text{o}}{2 \text{ mole h}_2\text{s}} \times 18 \text{ g h}_2\text{o} = 149.4 \text{ g h}_2\text{o}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g h}_2\text{o}}{149.4 \text{ g h}_2\text{o}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of o<sub>2</sub> produces a 47.3% yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have  $0.216 \text{ mole o}_2 \times \frac{32 \text{ g}}{1 \text{ mole}} = 6.91 \text{ g}$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{al}_2\text{o}_3 + 4 \text{al} + 3 \text{o}_2 \rightarrow 2 \text{al}_2\text{o}_3$  b determine the theoretical yield of  $\text{al}_2\text{o}_3$  160.0 g  $\text{al}$  x 1 mol  $\text{al}$  x 2 mol  $\text{al}_2\text{o}_3$  x 101.964 g  $\text{al}_2\text{o}_3$  302.3 g  $\text{al}_2\text{o}_3$  26.982 g  $\text{al}$  4 mol  $\text{al}$  1 mol  $\text{al}_2\text{o}_3$  c determine the percent yield  $\text{yield actual} \times 100$  260.0 g x 100 86.01 theoretical 302.3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of  $\text{h}_2$  are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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